

10th Science – Part1

Important Questions PDF (2026 SSC Board Exam)

Maharashtra State Board – Class 10

Chapterwise Most Expected Questions

Prepared by: (www.aplaabhyas.com)

★ **Note:** तुमच्या SSC परीक्षेस शुभेच्छा ✨



Chapter 1 .— Gravitation

✓ *Most Expected Questions & Answers for SSC 2026*

? Q1) State Newton's Universal Law of Gravitation.

✓ **Answer:**

Every object in the universe attracts every other object with a force. This force is directly proportional to the product of their masses and inversely proportional to the square of the distance between them.

$$F = \frac{GMm}{r^2}$$

? Q2) Define Universal Gravitational Constant (G) and write its SI unit.

✓ **Answer:**

Universal Gravitational Constant (G) is the constant of proportionality used in the universal law of gravitation.

SI Unit: $\text{N} \cdot \text{m}^2/\text{kg}^2$

? Q3) What is gravitational force? Explain with an example.

✓ **Answer:**

The attractive force acting between any two objects in the universe due to their masses is called **gravitational force**.

◆ Example: Earth pulls objects towards itself. Because of this gravitational force, when we throw a ball upward, it finally comes back to the ground.

? Q4) Explain the concept of free fall.

✓ **Answer:**

When an object falls freely towards the Earth under the influence of **only gravitational force**, without any air resistance or external force, it is called **free fall**.

? **Q5) What is acceleration due to gravity (g)? Give its value.**

✓ **Answer:**

The acceleration produced in a freely falling body due to the gravitational force of the Earth is called **acceleration due to gravity**.

Symbol: g

Average value on Earth: 9.8 m/s²

? **Q6) Define mass and weight. Write differences.**

✓ **Answer:**

Mass	Weight
Amount of matter in a body	Force with which Earth attracts a body
Constant everywhere	Changes with gravity
SI unit – kg	SI unit – N
Scalar quantity	Vector quantity

Relation:

$$W = mg$$

? **Q7) What is escape velocity? Give its formula.**

✓ **Answer:**

The minimum velocity required to escape from Earth's gravitational pull without falling back is called **escape velocity**.

$$v_e = \sqrt{2gR}$$

? **Q8) Solve the numerical:**

A body of mass 10 kg is lifted to a height of 8 m. Calculate the potential energy. ($g = 9.8 \text{ m/s}^2$)

✓ **Solution:**

Given:

$$m = 10 \text{ kg}$$

$$h = 8 \text{ m}$$

$$g = 9.8 \text{ m/s}^2$$

$$PE = mgh \quad PE = 10 \times 9.8 \times 8 = 784 \text{ J} \quad PE = 10 \times 9.8 \times 8 = 784 \text{ J}$$

✓ **Final Answer:**

The potential energy = **784 Joules**

✦ Chapter 1 — Gravitation (Additional Questions & Answers)

? **Q9) State the relation between gravitational force and distance between two bodies.**

✓ **Answer:**

Gravitational force is **inversely proportional to the square of the distance** between two bodies.

$$F \propto \frac{1}{r^2} \quad F \propto r^{-2}$$

Therefore, if distance is doubled, gravitational force becomes **one-fourth**.

? **Q10) State the relation between gravitational force and masses of two bodies.**

✓ **Answer:**

Gravitational force is **directly proportional to the product of the masses** of two bodies.

$$F \propto M \times m \quad F \propto M \times m$$

If mass increases → gravitational force also increases.

? **Q11) What is orbital velocity?**

✓ **Answer:**

The minimum velocity required for a satellite to revolve in a fixed orbit around Earth is called **orbital velocity**.

Formula:

$$v_o = \sqrt{\frac{gR}{2}} \quad v_o = 2gR$$

? Q12) What is weightlessness?

✓ **Answer:**

The condition in which a body experiences **no gravitational force** or appears to have **no weight** is known as **weightlessness**.

Example: Astronauts inside a satellite experience weightlessness.

? Q13) Give an example where gravitational force acts between small objects.

✓ **Answer:**

Two pieces of chalk kept on a table also attract each other due to gravitational force — but the force is extremely small, so we cannot feel it.

? Q14) Why does a stone fall faster than a feather?

✓ **Answer:**

Both experience gravitational force equally, but **air resistance** slows the feather more. If there is **no air resistance**, both fall at the same speed.

? Q15) What is the value of acceleration due to gravity (g) on the Moon?

✓ **Answer:**

Acceleration due to gravity on the Moon is **1.63 m/s²**, which is almost **one-sixth of Earth's gravity**.

? Q16) Why is gravitational force considered a universal force?

✓ **Answer:**

Because gravitational force acts **between every object in the entire universe** — whether very large (planet & star) किंवा खूप लहान (दोन पुस्तके).

? Q17) Why do objects thrown upward come back to the Earth?

✓ **Answer:**

Earth's gravitational force pulls every object towards its center. Therefore, an object thrown upward loses velocity and falls back.

? Q18) Why is potential energy considered gravitational potential energy?

✓ **Answer:**

Because potential energy of a body raised above the ground is stored **due to the gravitational attraction of Earth.**

Formula:

$$PE = mgh$$

Chapter 2 .— Periodic Classification of Elements

✓ *Most Expected Questions & Answers for SSC 2026*

? Q1) State Mendeleev's Periodic Law.

✓ **Answer:**

The properties of elements are a periodic function of their atomic masses.

That means, when elements are arranged in increasing order of atomic masses, **elements with similar properties repeat at regular intervals.**

? Q2) What is Modern Periodic Law?

✓ **Answer:**

According to the Modern Periodic Law:

The properties of elements are a periodic function of their atomic numbers.

→ Due to this, the modern periodic table is based on *atomic number*, not *atomic mass*.

? Q3) Why are noble gases placed in Group 18?

✓ **Answer:**

Noble gases have **completely filled outermost electron shells**, due to which:

- They are chemically stable
 - They do not react easily
- Hence, they are placed in **Group 18.**
-

? Q4) Define: (a) Period (b) Group

✓ Answer:

Term	Meaning
Period	A horizontal row of elements in the periodic table
Group	A vertical column of elements in the periodic table

? Q5) What is valency? Explain with an example.

✓ Answer:

The ability of an element to combine with other elements is called **valency**.

→ Example:

Sodium (Na) → atomic number 11 → electronic configuration 2,8,1 → valency = 1

? Q6) State any four physical properties of metals.

✓ Answer:

1. Good conductors of heat and electricity
 2. Malleable
 3. Ductile
 4. Have lustre (shiny appearance)
-

? Q7) Mention any four properties of non-metals.

✓ Answer:

1. Poor conductors of heat and electricity
 2. Non-malleable and non-ductile
 3. Brittle in nature
 4. Generally dull
-

? Q8) What are metalloids? Give examples.

✓ Answer:

Elements which show properties of **both metals and non-metals** are called *metalloids*.

Examples: Silicon (Si), Germanium (Ge)

? Q9) Explain the trend of atomic size:

- (a) Across a period
- (b) Down a group

✓ **Answer:**

Movement	Change in Atomic Size	Reason
Across a period	Decreases	Increase in nuclear charge
Down a group	Increases	Addition of new electron shells

? Q10) Give reasons:

✓ **(a) Sodium is stored in kerosene.**

Sodium reacts vigorously with oxygen and moisture in air. To prevent this reaction, it is stored in kerosene.

✓ **(b) Noble gases are chemically inert.**

Noble gases have a stable electronic configuration because their outermost shells are completely filled.

🔍 Chapter 3 — Chemical Reactions & Equations

✓ *Most Expected Questions & Answers for SSC 2026*

? Q1) What is a chemical reaction?

✓ **Answer:**

A chemical reaction is a process in which **one or more substances (reactants) are converted into new substances (products)** with different chemical properties.

? Q2) Write the balanced chemical equation for the following:

- (a) Magnesium + Oxygen → ?
(b) Iron + Copper Sulphate → ?

✓ **Answer:**

(a)



(b)

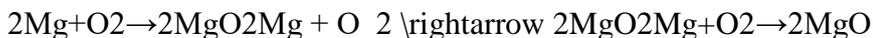


? Q3) What is oxidation?

✓ **Answer:**

Oxidation is a chemical reaction in which **a substance gains oxygen or loses hydrogen.**

Example:



? Q4) What is reduction?

✓ **Answer:**

Reduction is a chemical reaction in which **a substance loses oxygen or gains hydrogen.**

Example:



? Q5) What are displacement reactions? Give an example.

✓ **Answer:**

A displacement reaction is a reaction in which **a more reactive element displaces a less reactive element from its compound.**

Example:



? Q6) What are decomposition reactions?

✓ **Answer:**

A decomposition reaction is a reaction in which **a single compound breaks down into two or more simpler substances** when heated.

Example:



? Q7) What is corrosion? How can it be prevented?

✓ **Answer:**

Corrosion is the **slow degradation of metals due to reaction with moisture and oxygen.**

Prevention:

- Galvanization
 - Painting
 - Oiling and greasing
 - Electroplating
-

? Q8) What is rancidity?

✓ **Answer:**

Rancidity is a process in which **oily or fatty food items get oxidized and develop bad smell and taste.**
Example: Chips, ghee etc. become rancid if kept open for a long time.

? Q9) State two observations that indicate a chemical reaction has taken place.

✓ **Answer:**

1. Change in colour
 2. Formation of gas
 3. Change in temperature
 4. Formation of precipitate (any two)
-

? Q10) Why should chemical equations be balanced?

✓ **Answer:**

A chemical equation must be balanced because it follows the **Law of Conservation of Mass**, which states that:

“Mass can neither be created nor destroyed during a chemical reaction.”

Chapter 4. — Carbon Compounds

✓ *Most Expected Questions & Answers for SSC 2026*

? Q1) Why does carbon form a large number of compounds?

✓ **Answer:**

Carbon forms a large number of compounds due to:

1. **Catenation** (Ability to form long chains of carbon atoms)
 2. **Tetravalency** (Valency = 4 → can form 4 bonds)
 3. **Formation of single, double and triple bonds**
 4. **Stable covalent bonds**
-

? Q2) What is a covalent bond?

✓ Answer:

A covalent bond is a chemical bond formed by **mutual sharing of electrons between two atoms**.

Example: Methane (CH₄)

? Q3) What is homologous series?

✓ Answer:

A homologous series is a group of organic compounds having:

- Same functional group
- Same general formula
- Similar chemical properties
- Successive members differ by **-CH₂ (14 amu)**

Example: Alkane series — CH₄, C₂H₆, C₃H₈...

? Q4) Define isomerism.

✓ Answer:

Isomerism is the phenomenon in which **compounds have the same molecular formula but different structural arrangement of atoms**.

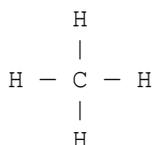
Example: C₄H₁₀ → n-butane and iso-butane

? Q5) Draw the structural formula of the following:

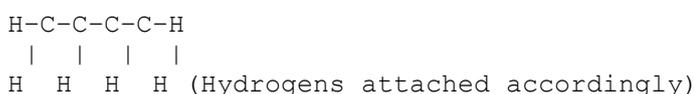
- (a) Methane
(b) Butane

✓ Answer:

- (a) Methane (CH₄)



- (b) Butane (C₄H₁₀)



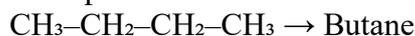
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? Q6) Explain the nomenclature of carbon compounds (IUPAC rules) in brief.

✓ **Answer:**

1. Select the longest carbon chain → root name
2. Number the carbon chain
3. Identify substituents
4. Write substituent name + root name + suffix

Example:



? Q7) What is ethanol? State any two uses.

✓ **Answer:**

Ethanol ($\text{C}_2\text{H}_5\text{OH}$) is a **liquid organic compound** used as:

1. **Alcoholic beverages**
 2. **Industrial solvent / fuel additive**
-

? Q8) What is ethanoic acid (acetic acid)? Write its common uses.

✓ **Answer:**

Ethanoic acid (CH_3COOH) is a weak organic acid used in:

1. Vinegar (Preservative)
 2. Pickles and food industry
-

? Q9) Why is carbon tetrachloride (CCl_4) non-inflammable?

✓ **Answer:**

Carbon tetrachloride does not burn because it **does not support combustion** and inhibits the burning process.

? Q10) Why are covalent compounds generally poor conductors of electricity?

✓ **Answer:**

Covalent compounds do not contain **free ions or charged particles**, so they **do not conduct electricity**.

🔑 Chapter 5.— Heat

✓ *Most Expected Questions & Answers for SSC 2026*

? Q1) What is specific heat capacity?

✓ **Answer:**

Specific heat capacity is the amount of heat energy required to raise the temperature of **1 kg of a substance by 1°C**.

Symbol: c

SI Unit: J/kg°C

? Q2) Define latent heat.

✓ **Answer:**

Latent heat is the amount of heat energy required to change the **state of a substance without any change in temperature**.

Types:

1. Latent heat of fusion
 2. Latent heat of vaporization
-

? Q3) What is heat conduction?

✓ **Answer:**

Heat conduction is the process of transfer of heat **from a region of higher temperature to a region of lower temperature** without actual movement of particles.

Example: Heating of metal rod from one end.

? Q4) What is greenhouse effect?

✓ **Answer:**

The trapping of heat in the Earth's atmosphere due to **greenhouse gases such as CO₂, methane and water vapour** is called the **greenhouse effect**.

Effect: Causes rise in global temperature.

? Q5) State the laws of heat conduction.

✓ **Answer:**

1. Heat flows from hot body to cold body.
 2. Rate of heat transfer is directly proportional to the area of cross-section and temperature difference.
 3. Rate of heat transfer is inversely proportional to the length of the conductor.
-

? Q6) Why are woolen clothes preferred in winter?

✓ **Answer:**

Wool contains **air trapped between its fibres**, which is a **bad conductor of heat**. Therefore, woolen clothes **prevent heat loss** from the body and keep us warm.

? Q7) Why are cooking utensils made of metals but their handles are made of plastic or wood?

✓ **Answer:**

Metals are **good conductors of heat** — they heat quickly while cooking. Plastic/wood are **bad conductors (insulators)** — so the handle does not become hot and can be held safely.

? Q8) What is anomalous expansion of water?

✓ **Answer:**

Water **contracts on cooling from 100°C to 4°C**, but **expands on cooling from 4°C to 0°C**. This unusual behaviour is called **anomalous expansion of water**.

? Q9) Define thermal conductivity.

✓ **Answer:**

Thermal conductivity is the ability of a material to **conduct heat**. Material with higher thermal conductivity → better conductor of heat.

? Q10) Solve the numerical:

How much heat is required to increase the temperature of 500 g of water from 25°C to 65°C?
(Specific heat capacity of water = 4.18 J/g°C)

✓ **Solution:**

Given:

$$m = 500 \text{ g}$$

$$c = 4.18 \text{ J/g}^\circ\text{C}$$

$$\Delta T = 65 - 25 = 40^\circ\text{C}$$

$$Q = mc\Delta T \quad Q = 500 \times 4.18 \times 40 = 83\,600 \text{ J}$$

✓ **Final Answer:**

Heat required = **83,600 Joules**

🔍 Chapter 6 — Refraction of Light

✓ *Most Expected Questions & Answers for SSC 2026*

? **Q1) What is refraction of light?**

✓ **Answer:**

Refraction of light is the **bending of light when it travels from one transparent medium to another** due to change in speed of light.

? **Q2) State Snell's Law.**

✓ **Answer:**

According to Snell's Law:

$$\frac{\sin i}{\sin r} = n$$

where,

i = angle of incidence

r = angle of refraction

n = refractive index of second medium

? **Q3) What is refractive index?**

✓ **Answer:**

Refractive index of a medium is the **ratio of the speed of light in air (or vacuum) to the speed of light in that medium.**

$$n = \frac{c}{v} \quad n = \frac{c}{v}$$

? **Q4) Why does a coin at the bottom of a glass full of water appear raised?**

✓ **Answer:**

The coin appears raised because **light rays coming from the coin bend away from the normal when they pass from water to air**, making the coin appear at a higher position.

? **Q5) Why does a stick appear bent when it is partly immersed in water?**

✓ **Answer:**

Due to **refraction of light from water to air**, the light rays bend and reach the eyes, creating a **shifted image**, making the stick appear bent.

? **Q6) Define: (a) Angle of incidence (b) Angle of refraction**

✓ **Answer:**

Term	Definition
Angle of incidence	Angle between the incident ray and normal
Angle of refraction	Angle between the refracted ray and normal

? **Q7) What is lateral displacement?**

✓ **Answer:**

Lateral displacement is the **sideways shift of the emergent ray** when a light ray passes through a glass slab.

? **Q8) Why does refraction not occur when light enters a medium perpendicular to the surface?**

✓ **Answer:**

When light enters **perpendicularly**, angle of incidence = 0° and the ray **does not bend**, so refraction does not occur.

? Q9) Why is the bottom of a swimming pool seen to be less deep than its actual depth?

✓ **Answer:**

Due to refraction, light rays from the bottom bend when moving from water to air, making the depth appear less than actual.

? Q10) Explain the relation between speed of light and refractive index.

✓ **Answer:**

Refractive index is **inversely proportional to the speed of light in a medium.**

$$n \propto \frac{1}{v} \text{ or } v \propto \frac{1}{n}$$

◆ Higher refractive index → Lower speed of light

◆ Lower refractive index → Higher speed of light

Chapter 7 — Lenses

✓ *Most Expected Questions & Answers for SSC 2026*

? Q1) What is a lens? Name its types.

✓ **Answer:**

A lens is a **transparent medium bounded by two refracting surfaces.**

Types of lenses:

1. **Convex lens (Converging lens)**
 2. **Concave lens (Diverging lens)**
-

? Q2) Define: (a) Principal axis (b) Focal length

✓ **Answer:**

Term	Definition
Principal axis	The straight line passing through the optical centre and the centres of curvature of a lens
Focal length	The distance between the optical centre and the principal focus of a lens

? Q3) What is focal point (principal focus)?

✓ **Answer:**

The principal focus of a lens is a point on the principal axis where **light rays parallel to the principal axis converge (convex lens) or appear to diverge from (concave lens)** after refraction.

? **Q4) Draw ray diagram showing image formation by a convex lens when the object is placed beyond 2F.**

✓ **Answer:**

Image position → Between F and 2F on the other side

Properties → Real, inverted, diminished

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? **Q5) State the Lens Formula and Magnification Formula.**

✓ **Answer:**

$$\frac{1}{v} - \frac{1}{u} = \frac{1}{f}$$

Magnification formula:

$$m = \frac{v}{u}$$

or

$$m = \frac{h_2}{h_1}$$

? **Q6) What is power of a lens? Give its SI unit.**

✓ **Answer:**

Power of a lens is the **reciprocal of its focal length in metre.**

$$P = \frac{1}{f}$$

SI unit: Dioptre (D)

? **Q7) Why is a concave lens used to correct myopia?**

✓ **Answer:**

In myopia (near-sightedness), **distant objects appear blurred because the image forms in front of the retina.**

A concave lens **diverges light rays** so that the image shifts backward and forms on the retina, helping to see distant objects clearly.

? Q8) List the nature of images produced by concave lens.

✓ Answer:

Concave lens always forms images that are:

1. Virtual
 2. Erect
 3. Diminished
 4. On the same side of the lens as the object
-

? Q9) What is optical centre of a lens?

✓ Answer:

The optical centre is a point inside the lens through which **light passes without deviation or without refraction.**

? Q10) Why do convex lenses burn paper in sunlight?

✓ Answer:

The convex lens **converges parallel rays from the Sun to a single point (focus)**, producing heat energy sufficient to burn paper.

Chapter 8 — Metallurgy

✓ Most Expected Questions & Answers for SSC 2026

? Q1) What is metallurgy?

✓ Answer:

Metallurgy is the branch of science that deals with **extraction of metals from their ores and making them suitable for practical use.**

? Q2) Define: (a) Ore (b) Gangue

✓ Answer:

Term**Definition**

Ore A naturally occurring mineral from which metal can be extracted economically
Gangue Impurities present in an ore such as sand, clay, etc.

? Q3) What is concentration of ore?**✓ Answer:**

The process of **removing impurities (gangue) from the ore** is called concentration of ore.

? Q4) Explain froth flotation process.**✓ Answer:**

Froth flotation is a method used to separate **sulphide ores** from impurities.

Steps:

1. Ore is mixed with water and oil and air is blown
 2. Sulphide ores stick to oil and float as froth
 3. Impurities settle down at the bottom
-

? Q5) What is roasting?**✓ Answer:**

Roasting is a process in which **sulphide ores are heated in presence of excess air** to convert them into oxides.

? Q6) What is smelting?**✓ Answer:**

Smelting is the process of **heating the ore with a reducing agent at high temperature** to extract metal in its molten form.

? Q7) What is alloy? Give two examples.**✓ Answer:**

An alloy is a **homogeneous mixture of two or more metals, or a metal and a non-metal**.

Examples:

- Brass (Copper + Zinc)
- Stainless steel (Iron + Chromium + Nickel)

? Q8) What is corrosion? State one method to prevent corrosion.

✓ Answer:

Corrosion is the **slow destruction of metals due to reaction with air and moisture.**

Prevention: Galvanization, painting, greasing, electroplating etc.

? Q9) Differentiate between metals and non-metals (any four).

✓ Answer:

Metals	Non-metals
Good conductors of heat & electricity	Poor conductors
Malleable	Non-malleable
Ductile	Non-ductile
Lustrous	Dull appearance

? Q10) Name the following:

✓ Answer:

Question	Answer
A metal that is liquid at room temperature	Mercury
Most reactive metal	Potassium
A metal that does not corrode	Gold
A lightweight metal used in aircraft	Aluminium

Chapter 9 — Space Mission

✓ Most Expected Questions & Answers for SSC 2026

? Q1) What is a space mission?

✓ **Answer:**

A space mission is a **scientific and technological program designed to send spacecraft, satellites or astronauts into space** for research and exploration.

? **Q2) What is satellite?**

✓ **Answer:**

A satellite is an **object that revolves around a planet in a fixed orbit.**

Satellites are of two types:

1. **Natural satellite** – e.g., Moon
 2. **Artificial satellite** – Man-made satellites launched into orbit
-

? **Q3) Explain remote sensing.**

✓ **Answer:**

Remote sensing is the technique of **collecting information about the Earth's surface from a distance using satellites** without physical contact.

Uses: Agriculture, weather forecasting, map-making, disaster monitoring

? **Q4) What is telemetry system in satellites?**

✓ **Answer:**

Telemetry is a system in a satellite that **collects, measures, and sends important data to Earth** such as temperature, pressure and system performance.

? **Q5) What is rocket propulsion?**

✓ **Answer:**

Rocket propulsion is the **motion of a rocket due to the backward expulsion of high-speed gases**, which creates a forward reaction force (Newton's 3rd law).

? **Q6) Explain the need of heat shield in spacecraft.**

✓ **Answer:**

When spacecraft returns to Earth, **air friction increases temperature to extremely high levels.**

A heat shield protects the spacecraft and astronauts from **burning due to heat.**

? Q7) State the uses of artificial satellites.

✓ **Answer:**

1. Communication
 2. Weather forecasting
 3. GPS and navigation
 4. Remote sensing and resource mapping
 5. Scientific research
-

? Q8) Why is the launch of rocket carried out from the Earth's equatorial region?

✓ **Answer:**

Due to the **rotation of the Earth**, the equatorial region provides **maximum rotational speed (highest velocity)** which helps rockets gain extra speed and reduce fuel consumption.

? Q9) What is the significance of ISRO for India?

✓ **Answer:**

ISRO strengthens India by:

- Developing satellites at low cost
 - Launching missions like Chandrayaan, Mangalyaan
 - Improving communication, weather forecasting, education & defence
 - Creating global recognition in space science
-

? Q10) Why does a satellite not fall to the Earth?

✓ **Answer:**

A satellite continuously moves **forward with a high tangential velocity** while also being pulled towards the Earth by gravity.

These two effects balance each other, so the satellite **remains in orbit and does not fall**.

Chapter 10 — Disaster Management

✓ *Most Expected Questions & Answers for SSC 2026*

? Q1) What is disaster?

✓ **Answer:**

A disaster is a **sudden and harmful event** that causes **loss of life, property or environment** and disrupts normal functioning of society.

? Q2) Define disaster management.

✓ **Answer:**

Disaster management refers to the **systematic approach of planning, organizing and responding to disasters** to minimize loss and help in quick recovery.

? Q3) State the types of disasters with examples.

✓ **Answer:**

Type	Examples
Natural disasters	Earthquake, flood, tsunami, cyclone
Man-made disasters	Fire, nuclear accident, terrorist attack, chemical explosion

? Q4) What is mock drill?

✓ **Answer:**

A mock drill is a **practice exercise to check the preparedness of people and systems during emergency situations**.

It trains people to act quickly and safely during disasters.

? Q5) What is the role of first responder in disaster?

✓ **Answer:**

The first responder provides **immediate help to victims before professional emergency services arrive**. Duties include:

1. Calling emergency services
 2. Giving first aid
 3. Helping in evacuation
 4. Maintaining safety
-

? Q6) Explain the disaster management cycle.

✓ **Answer:**

Disaster management cycle includes:

Stage	Explanation
Prevention	Avoiding the occurrence of disaster
Mitigation	Reducing the impact
Preparedness	Planning and readiness
Response	Immediate action during disaster
Recovery	Returning to normal situation

? Q7) Why is awareness important in disaster management?

✓ **Answer:**

Awareness helps people to:

- Understand danger
 - Take preventive actions
 - Reduce injuries and property loss
 - Act safely during emergencies
-

? Q8) Why is communication important during disasters?

✓ **Answer:**

Communication helps to:

- Share warnings and information quickly
 - Coordinate rescue teams
 - Provide correct guidance to public
 - Avoid panic and confusion
-

? Q9) What are the four priority actions of Sendai Framework?

✓ **Answer:**

1. Understanding disaster risk
 2. Strengthening disaster risk governance
 3. Investing in risk reduction
 4. Enhancing disaster preparedness and recovery
-

? Q10) State any four emergency numbers useful during disasters.

✓ **Answer:**

Service	Number
Police	100
Fire brigade	101
Ambulance	108 / 102
Disaster helpline	1077

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